**Stoichiometry Unit Objectives**

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|  | Objectives/Questions | Assignments |
|  | Review:1. What is a mole? Why do we use moles?
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|  | 1. What is Avogadro’s number? How is it used?
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|  | 1. What is molar mass? How is molar mass calculated?
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|  | 1. How do we convert from moles to mass to particles?
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|  | 1. What information do coefficients provide in a chemical formula?
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|  | New:1. Describe the difference between composition stoichiometry and reaction stoichiometry.
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|  | 1. Why do use stoichiometry?
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|  | 1. What is a mole ratio and why is it useful?
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|  | 1. Diagram how to convert from mass to moles to calculate each of unknowns.
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|  | 1. Make conversions involving:
	1. Moles to moles
	2. Moles to mass
	3. Mass to moles
	4. Mass to mass
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|  | 1. Describe the difference between a limiting and an excess reagent. (A diagram may be helpful)
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|  | 1. Identify the limiting reactant in a chemical equation.
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|  | 1. Identify the amount of excess reactant leftover.
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|  | 1. Calculate the amount of product produced.
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|  | 1. Distinguish between theoretical yield, actual yield, and percentage yield.
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|  | 1. Calculate percent yield given actual and theoretical yield.
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**Stoichiometry Vocabulary**

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| Word | Definition | Sentence or Picture |
| 1. **Actual yield**
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| 1. **Excess reagent**
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| 1. **Limiting reagent**
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| 1. **Mole ratio**
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| 1. **Percent yield**
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| 1. **Stoichiometry**
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| 1. **Theoretical yield**
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