**Stoichiometry Unit Objectives**

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|  | Objectives/Questions | Assignments |
|  | Review:   1. What is a mole? Why do we use moles? |  |
|  | 1. What is Avogadro’s number? How is it used? |  |
|  | 1. What is molar mass? How is molar mass calculated? |  |
|  | 1. How do we convert from moles to mass to particles? |  |
|  | 1. What information do coefficients provide in a chemical formula? |  |
|  | New:   1. Describe the difference between composition stoichiometry and reaction stoichiometry. |  |
|  | 1. Why do use stoichiometry? |  |
|  | 1. What is a mole ratio and why is it useful? |  |
|  | 1. Diagram how to convert from mass to moles to calculate each of unknowns. |  |
|  | 1. Make conversions involving:    1. Moles to moles    2. Moles to mass    3. Mass to moles    4. Mass to mass |  |
|  | 1. Describe the difference between a limiting and an excess reagent. (A diagram may be helpful) |  |
|  | 1. Identify the limiting reactant in a chemical equation. |  |
|  | 1. Identify the amount of excess reactant leftover. |  |
|  | 1. Calculate the amount of product produced. |  |
|  | 1. Distinguish between theoretical yield, actual yield, and percentage yield. |  |
|  | 1. Calculate percent yield given actual and theoretical yield. |  |

**Stoichiometry Vocabulary**

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| Word | Definition | Sentence or Picture |
| 1. **Actual yield** |  |  |
| 1. **Excess reagent** |  |  |
| 1. **Limiting reagent** |  |  |
| 1. **Mole ratio** |  |  |
| 1. **Percent yield** |  |  |
| 1. **Stoichiometry** |  |  |
| 1. **Theoretical yield** |  |  |